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As liquids absorb heat at their boiling
points, the temperature remains
constant while they vaporize. true

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(continued) Use the heating curve for
water shown below to answer
Questions 5, 6, and 7.

section 17.1 the flow of energy heat

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Complete the enthalpy diagram for
the combustion of natural gas. Use
the thermochemical equation in the
first paragraph on page 517 as a
guide. SECTION 17.3 HEAT IN
CHANGES OF STATE (pages 520–526)
This section explains heat transfers
that occur during melting, freezing,
boiling, and condensing.

SECTION 17.1 THE FLOW OF ENERGY
HEAT AND WORK (pages 505–510)
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1. How much energy must be absorbed by 20.0 g of water to increase its temperature from 283.0 ° C to 303.0 ° C? 2. When 15.0 g of steam drops in temperature from 275.0 ° C to 250.0 ° C, how much heat energy is released?

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